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### (54) Introducing treating fluids into subterranean producing zones

(57) Treating fluids are uniformly introduced into a subterranean producing zone by packing perforation tunnels in the producing zone with a particulate solid anhydrous borate material which degrades and dissolves in water over time. A treating fluid is then introduced into the subterranean zone by way of the packed perforation tunnels whereby the treating fluid uniformly flows into

the producing zone. The particulate solid anhydrous borate material in the perforation tunnels is allowed to degrade and dissolve in an aqueous fluid in contact therewith so that the aqueous fluid and produced formation fluids freely flow into the well bore.

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**Description**

[0001] The present invention concerns a method of uniformly introducing treating fluids into subterranean producing zones having perforation tunnels formed therein.

[0002] After a well bore is drilled and completed in a subterranean producing zone including the formation of perforation tunnels through casing and cement into the zone, it is often necessary to introduce a treating fluid into the zone. For example, the producing zone can be stimulated by introducing an aqueous acid solution to thereby increase its porosity and the production of hydrocarbons therefrom. In order to ensure that the producing zone is contacted by the treating fluid uniformly, it has heretofore been the practice to pack the perforation tunnels in the subterranean zone with a particulate solid diverting agent. The presence of the diverting agent in the perforation tunnels causes the treating fluid introduced therein to be uniformly distributed between all of the perforations whereby the subterranean zone is uniformly treated.

[0003] A problem in the use of the above described procedure is that the particulate solid diverting agent must subsequently be removed from the perforation tunnels. This has been accomplished by contacting the particulate solid diverting agent with a fluid which dissolves the diverting agent, i.e. water, oil, xylene, etc. While the clean-up procedure required can be successful in removing the diverting agent, it often entails considerable time and expense to carry out.

[0004] Thus, there is a need for an improved method of uniformly introducing a treating fluid into a subterranean producing zone by way of perforation tunnels therein involving simpler subsequent removal of the particulate solid diverting agent.

[0005] In one aspect, the present invention provides a method of uniformly introducing a treating fluid into a subterranean producing zone penetrated by a well bore having perforation tunnels formed therein, which method comprises the steps of:

- (a) packing the perforation tunnels with a particulate solid anhydrous borate material which degrades and dissolves in water over time;
- (b) introducing a treating fluid into said subterranean zone by way of said perforation tunnels containing said particulate solid anhydrous borate material, whereby said treating fluid uniformly flows into said producing zone; and
- (c) allowing said particulate solid anhydrous borate material in said perforation tunnels to degrade and dissolve in an aqueous fluid in contact therewith so that said aqueous fluid and produced formation fluids freely flow through said perforation tunnels into said well bore.

[0006] In the method of the present invention, a par-

ticulate solid anhydrous borate material which degrades and dissolves in water over time is packed into the perforation tunnels and into the subterranean producing zone. Thereafter, a treating fluid is introduced into the

5 subterranean zone by way of the perforation tunnels containing the particulate solid anhydrous borate material whereby the treating fluid uniformly flows into the producing zone. The particulate solid anhydrous borate material in the perforation tunnels is then allowed to de-  
10 grade and dissolve in an aqueous fluid in contact therewith so that the aqueous fluid and produced formation fluids freely flow through the perforation tunnels into the well bore.

[0007] The particulate solid anhydrous borate materials which can be utilized in accordance with this invention include, but are not limited to, anhydrous sodium tetraborate (also known as anhydrous borax), anhydrous boric acid and anhydrous boric oxide. These anhydrous borate materials are only slightly soluble in water. However, with time and heat in the subterranean zone, the anhydrous borate materials react with the surrounding aqueous fluid and are hydrated. The resulting hydrated borate materials are highly soluble in water as compared to the anhydrous borate materials and as a  
20 25 result are dissolved in the aqueous fluid. The total time required for the anhydrous borate materials to degrade and dissolve in an aqueous fluid is in the range of from about 8 hours to about 72 hours depending upon the temperature of the subterranean zone in which they are placed.

[0008] The particulate solid anhydrous borate material utilized is packed into the perforation tunnels in the subterranean zone by introducing a carrier fluid containing the particulate solid anhydrous borate material into the subterranean zone. The carrier fluid dissipates into the subterranean zone by way of the perforation tunnels and the particulate solid anhydrous borate material is screened out of the carrier fluid by the formation whereby it is packed into the perforation tunnels. A variety of carrier fluids can be utilized for packing the perforation tunnels with the anhydrous borate material utilized including, but not limited to, water, brines, seawater and formation water. Of these, brines and seawater are preferred.

[0009] As mentioned, the treating fluid which is uniformly introduced into the subterranean producing zone can be any of a variety of treating fluids, including, but not limited to, oil, oil-water emulsions, oil base foams, water base foams, fresh water, salt water, formation water and various aqueous solutions. Examples of aqueous treating solutions include, but are not limited to, aqueous acid solutions, aqueous scale inhibiting material solutions, aqueous water blocking material solutions, aqueous clay stabilizer solutions, aqueous chelating agent solutions, aqueous surfactant solutions and aqueous paraffin removal solutions.

[0010] When the aqueous treating fluid is an aqueous acid solution, the aqueous acid solution can include one

or more acids such as hydrochloric acid, hydrofluoric acid, acetic acid, formic acid and other organic acids. In acidizing procedures for increasing the porosity of subterranean producing zones, a mixture of hydrochloric and hydrofluoric acids is commonly utilized.

[0011] Another aqueous treating fluid which can be uniformly introduced into the subterranean producing zone in accordance with this invention is a solution of an aqueous scale inhibitor material. The aqueous scale inhibitor material solution can contain one or more scale inhibitor materials including, but not limited to, tetrasodium ethylenediamine acetate, pentamethylene phosphonate, hexamethylenediamine phosphonate and polyacrylate. These scale inhibitor materials attach themselves to the subterranean zone surfaces whereby they inhibit the formation of scale in tubular goods and the like when hydrocarbons and water are produced from the subterranean zone.

[0012] Another aqueous treating solution which can be utilized is a solution of an aqueous water blocking material. The water blocking material solution can contain one or more water blocking materials which attach themselves to the formation in water producing areas whereby the production of water is reduced or terminated. Examples of water blocking materials that can be used include, but are not limited to, sodium silicate gels, organic polymers with metal cross-linkers and organic polymers with organic cross-linkers. Of these, cross-linked organic polymers are preferred.

[0013] As will now be understood by those skilled in the art, any of a variety of treating fluids can be uniformly introduced into a subterranean producing zone having a perforation tunnel formed therein in accordance with this invention. Because the treating fluid is uniformly distributed into the subterranean producing zone, the treating fluid reaches all parts of the subterranean zone instead of only a portion of the zone. After the treating fluid has been placed in the zone, the particulate solid anhydrous borate material used degrades and dissolves in an aqueous fluid in contact therewith which allows the aqueous fluid and formation fluids to freely flow into the well bore. The aqueous fluid in which the particulate solid anhydrous borate material is dissolved can be various aqueous fluids, including, but not limited to, water from an aqueous treating fluid introduced into the subterranean zone, overflush water introduced into the zone after a treating fluid is introduced therein and formation water produced in the subterranean zone.

[0014] As mentioned, the particulate solid anhydrous borate material utilized degrades when in contact with the aqueous fluid and converts to the hydrated form of borate material. The hydrated borate material then dissolves in the aqueous fluid thereby eliminating the need for contacting the subterranean zone with one or more clean-up fluids. Another advantage of the particulate solid anhydrous borate materials of this invention is that the melting points of the materials are high, i.e., 1367°F for anhydrous sodium tetraborate and 840°F for anhy-

drous boric oxide, and as a result, the materials do not soften at high subterranean zone temperatures.

[0015] A preferred method of this invention for uniformly introducing an aqueous treating fluid into a subterranean producing zone penetrated by a well bore having perforation tunnels formed therein is comprised of the steps of: (a) packing the perforation tunnels with a particulate solid anhydrous borate material which degrades and dissolves in water over time, (b) introducing a treating fluid into the subterranean zone by way of the perforation tunnels containing the particulate solid anhydrous borate material whereby the treating fluid uniformly flows into the producing zone; and (c) allowing the particulate solid anhydrous borate material in the perforation tunnels to degrade and dissolve in an aqueous fluid in contact therewith so that the aqueous fluid and produced formation fluids freely flow through the perforation tunnels into the well bore.

[0016] Another preferred method of this invention for uniformly introducing an aqueous treating fluid into a subterranean producing zone penetrated by a well bore, having perforation tunnels formed therein is comprised of the steps of: (a) introducing a carrier fluid containing a particulate solid anhydrous borate material which degrades and dissolves in water over time into the perforation tunnels whereby the carrier fluid dissipates into the subterranean zone and the particulate solid anhydrous borate material is packed in the perforation tunnels; (b) introducing a treating fluid into the subterranean zone by way of the perforation tunnels containing the particulate solid anhydrous borate material whereby the treating fluid uniformly flows into the producing zone; and (c) allowing the particulate solid anhydrous borate material in the perforation tunnels to degrade and dissolve in an aqueous fluid in contact therewith so that the aqueous fluid and produced formation fluids freely flow through the perforation tunnels into the well bore.

[0017] In order to further illustrate the methods of this invention, the following example is given.

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#### EXAMPLE

[0018] The example below demonstrates the degradation over time of anhydrous sodium tetraborate and anhydrous boric acid in seawater solutions of scale inhibitors and 15% hydrochloric solutions. The amount of borate material and the volume of solutions used in the degradation experiments were chosen to simulate down hole conditions (i.e., perforation and well bore volumes). 45

[0019] The degradation experiments were carried out in a sealed cell equipped with a sight glass, pressurized with nitrogen to 200 psi and a temperature of 250°F. The degradation of the borate materials was measured by recording the change in volume of the borate materials over time.

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[0020] Table I below shows the percent volume of degradation over time of anhydrous boric oxide in a 10% phosphonate containing a scale inhibitor/seawater so-

lution. During the first minutes of the experiment there was an initial swelling of the boric oxide resulting in a volume increase. Over time, a gradual degradation of the boric oxide took place with complete degradation of the material occurring in approximately 52 hours.

TABLE I

Degradation of Anhydrous Boric Oxide	
Time (hrs.)	% Volume of Solid
0	100
0.05	144.44
0.083	166.67
0.25	161.11
0.42	161.11
0.92	155.56
1.5	144.44
1.67	138.89
2.17	133.33
2.67	133.33
4.08	133.33
4.75	127.78
5.75	127.78
6.83	125
7.33	125
10.83	122.22
13.5	111.11
21	88.89
23	88.89
24	88.89
26.75	77.78
28	77.78
30	72.22
31.25	66.67
34.25	55.56
45.75	22.22
48.25	11.11
51.75	0

[0020] Table II shows the percent volume of degradation over time of anhydrous boric oxide in a 10% methylenephosphonate containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the boric oxide resulting in a volume increase. Over time, a gradual degradation of the boric oxide took place with complete degradation of the material occurring in approximately 47 hours.

TABLE II

Degradation of Anhydrous Boric Acid	
Time (hrs.)	% Volume of Solid
0	100
0.083	118.18

TABLE II (continued)

Degradation of Anhydrous Boric Acid	
Time (hrs.)	% Volume of Solid
0.17	122.72
0.42	118.18
0.92	109.09
1.5	104.55
1.83	104.55
3.17	100
4.33	100
5.5	100
7.5	100
11.5	90.91
22	63.64
24	54.55
27.75	45.45
31.25	36.36
46.75	0

[0021] Table III shows the percent volume of degradation over time of anhydrous boric oxide in a 10% ammonium salt containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the boric oxide resulting in a volume increase. Over time, a gradual degradation of the boric oxide took place with complete degradation of the material occurring in approximately 46 hours.

TABLE III

Degradation of Anhydrous Boric Oxide	
Time (hrs.)	% Volume of Solid
0	100
0.083	200
0.167	210
0.33	195
0.4167	175
0.5	160
1	130
3.25	110
4.75	100
5.75	100
11.25	90
22.25	60
26.75	50
27.75	45
28.75	40
31	40
46	0

[0022] Table IV shows the percent volume of degradation over time of anhydrous boric oxide in a 10% copolymer containing a scale inhibitor/seawater solution.

During the first minutes of the experiment there was an initial swelling of the boric oxide resulting in a volume increase. Over time, a gradual degradation of the boric oxide took place with complete degradation of the material occurring in approximately 53 hours.

TABLE IV

Degradation of Anhydrous Boric Oxide	
Time (hrs.)	% Volume of Solid
0	100
0.083	137.5
0.25	156.25
0.67	143.75
0.83	137.5
4.08	150
5.33	150
6.33	150
8.33	150
23.33	100
28.33	93.75
30.33	75
32.33	62.5
47.33	25
53.33	0

[0023] Table V shows the percent volume of degradation over time of anhydrous boric oxide in a 10% neutralized phosphonate containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the boric oxide resulting in a volume increase. Over time, a gradual degradation of the boric oxide took place with complete degradation of the material occurring in approximately 70 hours.

TABLE V

Degradation of Anhydrous Boric Oxide	
Time (hrs.)	% Volume of Solid
0	100
0.083	138.89
0.133	155.56
0.167	166.67
0.25	161.11
0.75	150
1.08	144.44
2	133.33
2.92	127.78
5.42	122.22
7.58	122.22
22.58	100
25.58	94.44
27.58	77.78
31.58	66.67

TABLE V (continued)

Degradation of Anhydrous Boric Oxide	
Time (hrs.)	% Volume of Solid
5	46.58
	33.33
10	49.58
	27.78
	51.58
	27.78
	52.58
	22.22
	53.58
	22.22
	55.58
	11.11
	69.58
	0

[0024] Table VI shows the percent volume of degradation over time of anhydrous boric oxide in a 15% HCl solution. During the first minutes of the experiment there was an initial swelling of the boric oxide resulting in a volume increase. Over time, a gradual degradation of the boric oxide took place with complete degradation of the material occurring in approximately 60 hours.

TABLE VI

Degradation of Anhydrous Boric Acid	
Time (hrs.)	% Volume of Solid
25	0
	100
30	0.083
	133.33
	0.167
	166.67
	0.25
	177.78
35	0.47
	172.22
	0.67
	166.67
	2.41
	166.67
	3.92
	188.89
	5.5
	183.33
	6.17
	177.78
	12.91
	155.56
	23.17
	122.22
40	25.5
	111.11
	28.5
	100
45	29.833
	94.44
	38.25
	55.56
	50.75
	22.22
	59.75
	0

[0025] Table VII shows the percent volume of degradation over time of anhydrous sodium tetraborate in a 10% phosphonate containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the sodium tetraborate resulting in a volume increase. Over time, a gradual degradation of the sodium tetraborate took place with about 70% of the material degraded in approximately 76 hours.

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TABLE VII

Degradation of Anhydrous Sodium Tetraborate	
Time (hrs.)	% Volume of Solid
0	100
0.17	105.26
0.25	94.74
0.58	94.74
2.33	94.74
3.33	89.47
5.58	78.95
10.33	73.68
20.83	63.16
23.08	63.16
26.08	63.16
29.08	63.16
44.33	52.63
76.53	31.58

[0026] Table VIII shows the percent volume of degradation over time of anhydrous sodium tetraborate in a 10% methylenephosphonate containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the sodium tetraborate resulting in a volume increase. Over time, a gradual degradation of the sodium tetraborate took place with about 70% of the material degraded in approximately 54 hours.

TABLE VIII

Degradation of Anhydrous Sodium Tetraborate	
Time (hrs.)	% Volume of Solid
0	100
0.083	105.26
0.33	105.26
0.83	105.26
1.5	100
2.17	100
3.17	84.21
4.33	84.21
5.5	84.21
7.5	84.21
11.5	73.68
22	63.16
24	63.16
27.75	52.63
31.25	52.63
46.75	52.63
53.75	31.58

[0027] Table IX shows the percent volume of degradation over time of anhydrous sodium tetraborate in a

10% ammonium salt containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the sodium tetraborate resulting in a volume increase. Over time, a gradual degradation of the sodium tetraborate took place with about 60% of the material degraded in approximately 23 hours.

TABLE IX

Degradation of Anhydrous Sodium Tetraborate	
Time (hrs.)	% Volume of Solid
0	100
0.083	120
0.33	120
1	120
2	120
3.5	60
4.5	60
5.5	60
7.75	50
22.75	40

[0028] Table X shows the percent volume of degradation over time of anhydrous sodium tetraborate in a 10% copolymer containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the sodium tetraborate resulting in a volume increase. Over time, a gradual degradation of the sodium tetraborate took place with about 90% of the material degraded in approximately 70 hours.

TABLE X

Degradation of Anhydrous Sodium Tetraborate	
Time (hrs.)	% Volume of Solid
0	100
0.33	105.56
0.5	88.89
0.75	88.89
1.58	88.89
3.42	88.89
4.08	77.78
6.08	72.22
8.08	72.22
23.08	55.56
28.08	50
32.08	44.44
47.08	33.33
50.08	33.33
52.08	27.78
54.08	25
56.08	22.22
70.58	11.11

[0029] Table XI shows the percent volume of degradation over time of anhydrous sodium tetraborate in a 10% neutralized phosphonate containing a scale inhibitor/seawater solution. During the first minutes of the experiment there was an initial swelling of the sodium tetraborate resulting in a volume increase. Over time, a gradual degradation of the sodium tetraborate took place with about 75% of the material degraded in approximately 98 hours.

TABLE XI

Degradation of Anhydrous Sodium Tetraborate	
Time (hrs.)	% Volume of Solid
0	100
0.25	90
0.5	90
0.83	90
1.33	990
1.67	85
2.67	85
3.17	80
3.67	80
5.17	75
19.83	60
22.83	60
27.83	60
43.83	60
50.83	50
67.83	45
75.83	40
91.83	25
98.33	25

[0030] Table XII shows the percent volume of degradation over time of anhydrous sodium tetraborate in a 15% HCl solution. During the first minutes of the experiment there was an initial swelling of the sodium tetraborate resulting in a volume increase. Over time, a gradual degradation of the sodium tetraborate took place with complete degradation of the material in approximately 80 hours.

TABLE XII

Degradation of Anhydrous Sodium Tetraborate	
Time (hrs.)	% Volume of Solid
0	100
0.13	84.62
0.18	76.92
0.67	53.85
0.75	53.85
0.92	53.85
1.33	46.15

TABLE XII (continued)

Degradation of Anhydrous Sodium Tetraborate	
Time (hrs.)	% Volume of Solid
4.83	46.15
6.83	46.15
11.33	46.15
22.33	42.31
30.58	40.38
35.58	36.54
46.58	30.77
50.58	26.92
53.17	26.92
56.17	26.92
72.67	23.08
75.67	23.08
80.67	0

[0031] Thus, the present invention is well adapted to carry out the objects and attain the ends and advantages mentioned as well as those which are inherent therein. While numerous changes can be made by those skilled in the art, such changes are encompassed within the spirit of this invention as defined by the appended claims.

### 30 Claims

1. A method of uniformly introducing a treating fluid into a subterranean producing zone penetrated by a well bore having perforation tunnels formed therein, which method comprises the steps of:
  - (a) packing the perforation tunnels with a particulate solid anhydrous borate material which degrades and dissolves in water over time;
  - (b) introducing a treating fluid into said subterranean zone by way of said perforation tunnels containing said particulate solid anhydrous borate material, whereby said treating fluid uniformly flows into said producing zone; and
  - (c) allowing said particulate solid anhydrous borate material in said perforation tunnels to degrade and dissolve in an aqueous fluid in contact therewith so that said aqueous fluid and produced formation fluids freely flow through said perforation tunnels into said well bore.
2. A method according to claim 1, wherein said particulate solid anhydrous borate material is selected from anhydrous sodium tetraborate, anhydrous boric acid and anhydrous boric oxide.
3. A method according to claim 1 or 2, wherein said particulate solid anhydrous borate material is

packed in said perforation tunnels by introducing a carrier fluid containing said particulate solid anhydrous borate material into said perforation tunnels whereby said carrier fluid dissipates into said subterranean zone and said particulate solid anhydrous borate material is packed in said perforation tunnels. 5

4. A method according to claim 3, wherein said carrier fluid is selected from water, brines, seawater and formation water. 10

5. A method according to claim 1, 2, 3 or 4, wherein said aqueous fluid in contact with said anhydrous borate material in accordance with step (c) is treating fluid water, overflush water or formation water. 15

6. The method according to any of claims 1 to 5, wherein said treating fluid introduced in accordance with step (b) is selected from oil, oil-water emulsions, oil base foams, water base foams, fresh water, salt water, formation water and aqueous solutions. 20

7. A method according to claim 5, wherein said treating fluid is an aqueous solution selected from aqueous acid solutions, aqueous scale inhibitor material solutions, aqueous water blocking material solutions, aqueous clay stabilizer solutions, aqueous chelating agent solutions, aqueous surfactant solutions and aqueous paraffin removal solutions. 25

8. A method according to claim 7, wherein said treating fluid is an aqueous acid solution comprising one or more acids selected from hydrochloric acid, hydrofluoric acid, acetic acid, formic acid and other organic acids. 35

9. A method according to claim 8, wherein said acid solution comprises a mixture of hydrochloric acid and hydrofluoric acid. 40

10. A method according to claim 7, wherein said treating fluid is an aqueous scale inhibitor material solution comprising one or more scale inhibitor materials selected from tetrasodium ethylenediamine acetate, pentamethylene phosphonate, hexamethylene diamine phosphonate, and polyacrylate. 45

11. A method according to claim 10, wherein said aqueous scale inhibitor material solution comprises tetrasodium ethylenediamine acetate and pentamethylene phosphonate. 50

12. A method according to claim 7, wherein said treating fluid is an aqueous water blocking material solution comprising one or more water blocking materials selected from sodium silicate gels, organic pol- 55

ymers with metal cross-linkers and organic polymers with organic cross-linkers.



DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int.Cl.)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
A	US 3 677 343 A (SHOWALTER WILLIAM E) 18 July 1972 (1972-07-18) * column 2, line 34 - column 4, line 22 * -----	1-12	E21B43/27
A	US 3 707 194 A (SVALDI MARVIN A) 26 December 1972 (1972-12-26) * column 1, line 55 - column 3, line 60 * -----	1-12	
			TECHNICAL FIELDS SEARCHED (Int.Cl.)
			E21B
<p>The present search report has been drawn up for all claims</p>			
Place of search	Date of completion of the search	Examiner	
MUNICH	20 January 2004	Zimpfer, E	
CATEGORY OF CITED DOCUMENTS			
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ANNEX TO THE EUROPEAN SEARCH REPORT  
ON EUROPEAN PATENT APPLICATION NO.

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20-01-2004

Patent document cited in search report		Publication date		Patent family member(s)	Publication date
US 3677343	A	18-07-1972	CA	938558 A1	18-12-1973
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For more details about this annex : see Official Journal of the European Patent Office, No. 12/82